

Reg No.: _____

Name: _____

APJ ABDUL KALAM TECHNOLOGICAL UNIVERSITY
FIRST SEMESTER B.TECH DEGREE EXAMINATION, DECEMBER 2017

Course Code: CY100

Course Name: ENGINEERING CHEMISTRY

Max. Marks: 100

Duration: 3 Hours

PART A

Answer all questions, each question carries 2 marks

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| 1 | Calculate the fundamental vibrational frequency HCl molecule, if the value of force constant of the molecule is 483 Nm^{-1} . The atomic masses are $^1\text{H} = 1.673 \times 10^{-27} \text{ kg}$ and $^{35}\text{Cl} = 58.06 \times 10^{-27} \text{ kg}$. | 2 |
| 2 | Represent electrode reactions at different electrodes of a Li ⁻ ion cell during discharging. | 2 |
| 3 | Distinguish between retention factor and retention time in chromatography | 2 |
| 4 | What is ABS? How is it prepared? | 2 |
| 5 | Define a) Octane number and b) Cetane number. | 2 |
| 6 | Calculate the theoretical GCV of a petroleum fuel with composition C= 84%, H= 15% , O= 0.4%, N= 0.3% and S=0.3% | 2 |
| 7 | Give the principle of reverse osmosis? | 2 |
| 8 | Calculate the BOD of a water sample containing 75 mg of carbohydrate (CH ₂ O) per litre. | 2 |

PART B

Answer all questions, each question carries 3 marks

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| 9 | State and explain Beer- Lamberts law. Mention any two limitations of the law. | 3 |
| 10 | A zinc rod is dipped 0.3 M CuSO ₄ solution. Displacement reactions take place and allowed to attain equilibrium. Find the equilibrium constant of the reaction. [Given that $E^0_{\text{Cu}^{2+}/\text{Cu}} = +0.34\text{V}$ and $E^0_{\text{Zn}^{2+}/\text{Zn}} = -0.76\text{V}$. | 3 |
| 11 | Explain the visualisation techniques in TLC. | 3 |
| 12 | Explain the synthesis and applications of polypyrrole. | 3 |
| 13 | On burning 0.75g of fuel in a bomb calorimeter, the temperature of 2000g of water increases from 27.0 °C to 29.8 °C. The water equivalent of calorimeter and latent heat of steam are 385.0g and 587.0 cal/g respectively. If the fuel contains 0.9% hydrogen, calculate its gross and net calorific values. | 3 |
| 14 | Write a short note on biodiesel. | 3 |
| 15 | Illustrate break point chlorination with the help of suitable graph. Give any two advantages of break point chlorination. | 3 |
| 16 | Briefly explain the UASB process for sewage water treatment. | 3 |

PART C

Each question carries 10 marks.

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| 17 | a) Outline the principles of IR spectroscopy. | 5 |
| | b) How will you distinguish ethanol and dimethyl ether using NMR spectroscopy? | 5 |

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OR

- 18 a) Define chemical shift in NMR spectroscopy. Also explain the factors influencing chemical shift. 5
b) Discuss the possible electronic transitions in acetaldehyde. 5
- 19 a) Outline the setting up of a calomel electrode with a neat diagram. 5
b) How will you employ the calomel electrode for the determination of electrode potential of copper and zinc? 5
- OR
- 20 a) Explain the working and electrode reactions in H₂-O₂ fuel cells. 5
b) Give the procedure for potentiometric estimation of an alkali using a standard acid. 5
- 21 Explain the principles, instrumentation, working and applications of TGA. 10
- OR
- 22 Explain the principles, instrumentation, working and applications of HPLC. 10
- 23 a) Briefly outline chemical synthesis of nanoparticles. 5
b) What are conducting polymers? Briefly explain their classification. 5
- OR
- 24 a) Discuss the classification of nanomaterials. 5
b) Outline the preparation, properties and application of silicone rubbers. 5
- 25 Discuss the determination of calorific value of a fuel using a bomb calorimeter. 10
- OR
- 26 a) What are lubricants? How are they classified? 5
b) Briefly outline following properties of lubricants a) viscosity index b) flash and fire points. 5
- 27 Describe the EDTA method for determination of Hardness of water. 10
- OR
- 28 a) What is meant by COD? How it is determined? Explain its significance in sewage treatment. 5
b) A sample of water on analysis gave the following results: Ca²⁺ = 200 mg/L, Mg²⁺ = 180 mg/L, HCO₃⁻ = 360 mg/L, Na⁺ = 80 mg/L and Cl⁻ = 200 mg/L. Calculate the temporary, permanent and total hardness of the sample. 5
